

## SOLUBILITIES IN WATER

i — nearly insoluble ss — slightly soluble s — soluble d — decomposes n — not isolated	acetate	bromide	carbonate	chloride	chromate	hydroxide	iodide	nitrate	phosphate	sulfate	sulfide
Aluminum	ss	s	n	s	n	i	s	s	i	s	d
Ammonium	s	s	s	s	s	s	s	s	s	s	s
Barium	s	s	i	s	i	s	s	s	i	i	d
Calcium	s	s	i	s	s	ss	s	s	i	ss	d
Copper(II)	s	s	i	s	i	i	n	s	i	s	i
Iron(II)	s	s	i	s	n	i	s	s	i	s	i
Iron(III)	s	s	n	s	i	i	n	s	i	ss	d
Lead	s	ss	i	ss	i	i	ss	s	i	i	i
Magnesium	s	s	i	s	s	i	s	s	i	s	d
Mercury(I)	ss	i	i	i	ss	n	i	s	i	ss	i
Mercury(II)	s	ss	i	s	ss	i	i	s	i	d	i
Potassium	s	s	s	s	s	s	s	s	s	s	s
Silver	ss	i	i	i	ss	n	i	s	i	ss	i
Sodium	s	s	s	s	s	s	s	s	s	s	s
Zinc	s	s	i	s	s	i	s	s	i	s	i

**How to use this table.** The names of several metals that normally form positive ions (plus the ammonium ion,  $\text{NH}_4^+$ ) are listed vertically at the left of the table. Across the top of the table are listed the names of several nonmetals and clusters of atoms that normally form negative ions. To obtain information about the solubility of any compound included in the table, find the box where the row for the positive ion and the column for the negative ion meet. The solubility information is in that box.