

## Science Vocabulary #3<sup>1</sup>

**Atomic Mass** - The average relative mass of the isotopes of an element.

**Isotopes** - An atom of an element that has the same atomic number, but different atomic mass.

**Atomic Number** - The number of protons in the nucleus of an atom.

**Nucleus** - The central part of an atom that contains the protons and neutrons.

**Subatomic Particles** - Refers to parts of an atom, e.g. the protons, neutrons and electrons

**Neutron** - A subatomic particle that is electrically neutral, and is in the nucleus of an atom.

**Electron** - A subatomic particle that exists outside the nucleus and carries a negative charge.

**Proton** - A subatomic particle found in the nucleus of an atom, and carries a positive electrical charge, and a mass of about 1 amu, (atomic mass unit)

**Atomic mass unit** - A unit of mass equal to 1/12 the mass of a carbon -12 atom.

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<sup>1</sup> Foundations of College Chemistry, Morris Hein and Susan Arena, Brooks/Cole Publishing, 10<sup>th</sup> Edition, Glossary Appendix G.